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**Method for Rate**

**of Reaction** *Some*

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Chemistry chapter  
11 Chemical  
reactions answer  
key. coefficient. a  
whole number that  
appears before a  
formula in an  
equation. spectator  
ion. a particle not  
directly involved in  
a chemical  
reaction.  
combustion  
reaction. a reaction

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in which oxygen reacts with another substance, often producing light or heat. reactant.

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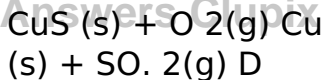


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in writing a  
balanced chemical  
equation. Write the  
skeleton equation  
for the following  
reactions: Heating  
solid  
copper(II)sulfide in  
the presence of  
oxygen gas  
produces pure  
copper and sulfur  
dioxide gas. Iron  
metal and chlorine

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gas react to form  
solid  
iron(III)chloride.



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PLAY. Match.

Gravity. Created  
by. tagreen348.

Terms in this set  
(51) (aq) dissolved  
in water (s) solid (l)  
liquid. reactants.  
the starting  
substances of a  
chemical reaction.  
products. the  
ending substances  
of a chemical  
reaction.

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11 - Chemical  
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Additional

Questions &

Answers. Question

1. Give an example

of a complex

chemical reaction.

Answer:

‘Photosynthesis’ is

a complex

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Chemical reaction.

Question 2. Give an example of simple chemical reactions.

Answer: Rusting of iron and combustion of fuels are simply chemical reactions.

Question 3.

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#### 11.2 Types of Chemical Reactions

Often charcoal briquettes provide the heat for barbeque grills through the burning of carbon. Have you ever felt the heat and smelled the smoke coming from a burning charcoal

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grill? The heat and smoke are the products of a combustion reaction.

Combustion is one of the five general types of chemical

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formulas for the products in these double replacement reactions.

1)  $\text{H}_3\text{PO}_4$   
2)  $\text{K}_2\text{CO}_3 + \text{BaCl}_2 \rightarrow$   
3)  $\text{Al} + \text{HNO}_3 \rightarrow$   
4)  $\text{AgNO}_3 + \text{KCl} \rightarrow$   
5)  $\text{H}_2\text{SO}_4 + \text{Co} \rightarrow$   
6)  $\text{BCl}_3 + \text{H}_2\text{O} \rightarrow$   
7)  $\text{P}_4\text{O}_{10} + \text{CaS} \rightarrow$   
8)  $\text{Ag}_2\text{CO}_3 \rightarrow$   
9)  $\text{Ag}_2\text{CO}_3 + \text{K}_2\text{CrO}_4 \rightarrow$

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Chemical Reactions

Answer Key

Chapter 11 ... 11.2

Types of Chemical

Reactions > 29

Whether one metal  
will displace

another metal from  
a compound

depends upon the  
relative reactivities

of the 11.2 Types  
of Chemical

Reactions > Section

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### 11.1 -Describing Chemical Reactions

The reactants are written on the left

and the products on the right. The

arrow that

separates them is called yield.

Chapter 11:

Chemical Reactions

## **Chapter 11 1**

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**Answer Key |**

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11.2 Types of  
Chemical

Reactions > 30 A halogen can also replace another halogen from a compound. • The activity of halogens decreases as you go down Group 7A of the periodic

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table—fluorine, chlorine, bromine, and iodine.

- Bromine is more active than iodine, so this reaction occurs:  $\text{Br}_2(\text{aq}) + 2\text{NaI}(\text{aq}) \rightarrow 2\text{NaBr}(\text{aq}) + \text{I}_2(\text{aq})$

## 11.2 Types of Chemical Reactions >

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Most chemical reactions can be classified into one or more of five

basic types:

acid-base

reactions: A

chemical reaction

that has the

general form  $AB +$

$C \rightarrow AC + B$  or  $AB +$

$CD \rightarrow AD + CB.$ ,

condensation

reactions: A

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Chemical reaction  
that has the  
general form  $A + B$   
 $\rightarrow AB$ .

Condensation  
reactions are the  
reverse of cleavage  
reactions.

## **Chapter 11.6: Types of Chemical Reactions - Chemistry ...**

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As shown in Figure  
11.3.1, applying a  
small amount of

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heat to a pile of orange ammonium dichromate powder results in a vigorous reaction known as the ammonium dichromate volcano. Heat, light, and gas are produced as a large pile of fluffy green chromium(III) oxide

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forms. We can describe this reaction with a chemical equation

An expression that gives the identities and quantities of the ...

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